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= 12.00 mmol - 5.40 mmol Unreacted ethanoic acid = 6.60 mmol Since 18.00 mL of NaOH(aq) was added to 20.00 mL of ethanoic acid solution, the total volume is now 38.00 mL. Solution: The concentration of unreacted ethanoic acid, [HC₂H₃O₂(aq)], can be determined: $\frac{23.2 \text{ HC}_2\text{H}_3\text{O}_2(\text{aq})}{23.2 \text{ [HC}_2\text{H}_3\text{O}_2(\text{aq})]} (6.60 \text{ mmol}) = 38.00 \text{ mL}$

Section 8.7: Acid-Base Titration Tutorial 1 Practice, page 547

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8 MHR Chemistry 12 Solutions Manual 978-0-07-106042-4 8. Identify any errors in the structure by drawing them. Rename the structure correctly. 2-ethyl-2,4,4-trimethylpentane What Is Required? You are to identify the errors in the name by drawing the structure and renaming it correctly. What Is Given? You are given an incorrectly named alkane.

Unit 1 Organic Chemistry - Mr. Arthur's Science Page

Tuesday August 12, 2014 . Review Questions, p. 39, #1-6. Define Carbonyl group, and aldehyde and ketone, Name IUPAC (Remember International Union of Pure and Applied Chemistry) Wednesday August 13, 2014 . Assessment on Naming of Alkanes, Alkenes, and Alkynes. Thursday August 14, 2014 . 1.6 Carboxylic Acids, Esters, and Fats . Define each one ...

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