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## **Chemistry 11 - kitsmini**

11. (a) The high/low solubility cutoff point is 0.10 mol/L at SATP. (b) The 0.10

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mol/L cutoff is useful because most laboratory solutions are within an order of magnitude of this

## **CHAPTER 7 SOLUBILITY AND REACTIONS**

11(s), (sucrose) because it dissolves well, to produce a noncon-ducting, neutral solution. Chemical 2 is the ionic

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compound,  $\text{KCl (s)}$ , because it dissolves well, to produce a conducting, neutral solution. Chemical 4 is the ionic hydroxide,  $\text{Ba(OH)}_2\text{(s)}$ , because it dissolves well, to produce a conducting, basic solution.

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