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25 terms. jakeydvd. chapter 12 study guide chemistry. law of conservation of

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mass. formula mass. law of definite proportions. limiting reactant. no detectable change in mass occurs during a chemical reaction. the sum of the atomic masses given in the formula of a substan....

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Change from Gas to Liquid and from
Liquid to Gas - An Introduction to
Dynamic Equilibrium The Process of
Condensation The Process of

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Evaporation How Evaporation Causes
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Chapter 12 - States of Matter Section
12.1 Gases 1. motion 2. a. small b.
forces c. random d. elastic; kinetic 3. $KE = \frac{1}{2}mv^2$ 4. Temperature 5. true 6. true
7. false 8. true 9. true 10. false 11. true
12. false 13. a 14. a 15. d 16. d 17. b 18.
b ...

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CHAPTER Section 12.4 Phase Changes In
your textbook, read about phase
changes. Date Class STUDY GUIDE
Complete the table by writing the initial
and final phases for each phase change
and making a check (V) in the correct
energy column. Phase Change Phase

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Answers

Energy released required initial final SO
Column B boiling point freezing point
melting point

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1. Chapter 12 - Chemical Kinetics. 12.1
Reaction Rates. A. Chemical kinetics 1.
Study of the speed with which reactants
are converted to products B. Reaction
Rate 1. The change in concentration of a
reactant or product per unit of time. t A t
 t concentration of A at time t
concentration of A at time t Rate. $2 \ 1 \ 2$
1.

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Matter and Change • Chapter 12 Section
12.2 Stoichiometric Calculations In your
textbook, read about mole-to-mole
conversion. Read the following passage
and then solve the problems. In the
equation that follows each problem,

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write in the space provided the mole ratio that can be used to solve the problem.

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